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Selective complexation of divalent cations by a cyclic α,β-peptoid hexamer: a spectroscopic and computational study
Selective complexation of divalent cations by a cyclic α,β-peptoid hexamer: a spectroscopic and computational study†

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We describe the qualitative and quantitative analysis of the complexation properties towards cations of a cyclic peptoid hexamer composed of alternating α- and β-peptoid monomers, which bear exclusively chiral (S)-phenylethyl side chains (spe) that have no noticeable chelating properties. The binding of a series of monovalent and divalent cations was assessed by 1H NMR, circular dichroism, fluorescence and molecular modelling. In contrast to previous studies on cations binding by 18-membered α-cyclopeptoid hexamers, the 21-membered cyclopeptoid cp1 did not complex monovalent cations (Na+, K+, Ag+) but showed selectivity for divalent cations (Ca2+, Ba2+, Sr2+ and Mg2+). Hexacoordinated C-3 symmetrical complexes were demonstrated for divalent cations with ionic radii around 1 Å (Ca2+ and Ba2+), while 5-coordination is preferred for divalent cations with larger (Ba2+) or smaller ionic radii (Mg2+).

In the realm of peptidomimetics, α-peptoids (poly-N-substituted glycines) are particularly attractive architectures that were previously shown to bind metal ions forming metallopeptoids.4 Peptoid synthesis is straightforward and an extremely high diversity of side chains is accessible.5 Their head to tail macrocyclisation is also particularly facile due to the inherent flexibility of the peptoid backbone.6 This also applies to the formation of cyclic β-peptoids7 and analogous cyclic alternated α,β-peptoids.8

Cyclic α-peptoids with 3,9 4, 6, 8 and 10 residues have been shown to bind cations, particularly from the first group alkali metals with selectivity depending on the ring size.10 For example, 18-membered cyclopeptoids showed a peak of selectivity for Na+ and metallated structures could be characterised in the solid state.10a,d These include the formation of the first supramolecular 1D metal–organic framework (MOF) based on a cyclic peptoid.10d In this MOF triggered by Na+, N-methoxethyl coordinating side chains participate to generate the supramolecular assembly. Recently, binding of Gd3+ cations was also demonstrated by cyclic α-hexapeptoids characterized by the presence of six N-carboxyethyl side chains or three N-methoxethyl and three carboxyethyl side chains in alternation.11 Interestingly, these side chains, which confer aqueous solubility to the complexes, do not appear to have any effective role in Gd3+ coordination.

Herein we present the first reported metal binding ability of a cyclic α,β-peptoid hexamer towards a selection of metal cations. The 21-membered cyclopeptoid cp1 is characterised by six chiral (S)-1-phenylethyl side chains (spe) that lack co-
Results and discussion

Synthesis

The cyclic α,β-peptoid cP1 was synthesised by head to tail cyclisation of its linear precursor using a previously described methodology. Briefly, the linear precursor was composed of (S)-N-(1-phenylethyl)glycine and (S)-N-(1-phenylethyl)-β-alanine monomers in alternation and was synthesised by a solution-phase submonomer approach. The C-terminal tBu-capping group of the synthesised linear peptoid was cleaved using trifluoroacetic acid and subsequent HATU-mediated cyclisation led to cP1 in 64% yield (see ESI† for details).

Experimental design

Metal binding is a delicate equilibrium between host–guest, host–solvent and guest–solvent interactions. Binding occurs when the interaction between the host and guest overcomes all of the other competing interactions. Thus, the experimental design is essential for the correct evaluation and quantification of the binding event. In this study acetonitrile (CH₃CN) was selected due to its low cation-coordinating properties. CH₃CN has also been extensively employed for conformational studies of α, β- and α,β-peptoids. Cyclic α-peptoids have already shown the ability to bind monovalent cations in CD₃CN by NMR. In addition, its transparency in the far and near regions of the UV spectrum makes it an ideal solvent for spectroscopic studies. Picrates were used as solvent interactions. Binding occurs upon addition of a molar excess of sodium or potassium picrates to cyclopeptoid cP1 (see ESI, Fig. S1†), indicating the absence of binding. Conversely, a significant overall simplification of the 1H NMR spectra was observed upon addition of a molar excess of sodium or potassium picrates to cyclopeptoid cP1 (see ESI, Fig. S2†), indicating that metal binding occurred. Binding of divalent cations (Sr²⁺, Ca²⁺, Mg²⁺ and Ba²⁺) was also observed when using perchlorate counterions, thus indicating that the counterion did not affect binding (Fig. 2).

Table 1 Metal ions investigated

<table>
<thead>
<tr>
<th>Cation</th>
<th>Ionic radius (Å)</th>
<th>Charge density (charge Å⁻²)</th>
<th>log Kₐ b</th>
</tr>
</thead>
<tbody>
<tr>
<td>+1 Na</td>
<td>1.02</td>
<td>0.076</td>
<td>0.91, 1.00^{15b}</td>
</tr>
<tr>
<td>+1 Ag</td>
<td>1.15</td>
<td>0.060</td>
<td>—</td>
</tr>
<tr>
<td>+1 K</td>
<td>1.38</td>
<td>0.042</td>
<td>1.23, 1.52^{15b,c}</td>
</tr>
<tr>
<td>+2 Mg</td>
<td>0.72</td>
<td>0.307</td>
<td>2.26, 2.37^{15d}</td>
</tr>
<tr>
<td>+2 Zn</td>
<td>0.74</td>
<td>0.291</td>
<td>1.681^{15e}</td>
</tr>
<tr>
<td>+2 Fe</td>
<td>0.78</td>
<td>0.262</td>
<td>—</td>
</tr>
<tr>
<td>+2 Ca</td>
<td>1.00</td>
<td>0.159</td>
<td>2.44, 2.74^{15d}</td>
</tr>
<tr>
<td>+2 Sr</td>
<td>1.18</td>
<td>0.114</td>
<td>2.58^{15d}</td>
</tr>
<tr>
<td>+2 Ba</td>
<td>1.35</td>
<td>0.087</td>
<td>2.69^{15d}</td>
</tr>
<tr>
<td>+3 Fe</td>
<td>0.55</td>
<td>0.789</td>
<td>—</td>
</tr>
</tbody>
</table>

a Ionic radius for coordination number VI. b log Kₐ values for M⁺ + nClO₄⁻ → Mⁿ⁺(ClO₄)ₙ in CH₃CN.

to assess whether cyclic α,β-peptoid cP1 displayed any selective metal binding ability among the cations listed in Table 1.

Metal ion complexation

1H NMR. The complexity of the 1H NMR spectra of cP1 in different solvents (acetone-d₆, CDCl₃ and CD₃CN) revealed several conformations in equilibrium on the NMR time scale, giving us the opportunity to assess the binding of cations by simplification of the spectra. The binding interaction between cP1 and a series of metal ions was therefore investigated by 1H NMR in CD₃CN. No spectral changes were observed upon addition of a molar excess of sodium or potassium picrates to cyclopeptoid cP1 (see ESI, Fig. S1†), indicating the absence of binding. Conversely, a significant overall simplification of the 1H NMR spectra was observed upon addition of strontium or calcium picrates, and to a lesser extent with magnesium picrates (see ESI, Fig. S2†), indicating that metal binding occurred. Binding of divalent cations (Sr²⁺, Ca²⁺, Mg²⁺ and Ba²⁺) was also observed when using perchlorate counterions, thus indicating that the counterion did not affect binding (Fig. 2).

Fig. 1 Chemical structure of 21-membered cyclic α,β-peptoid cP1.

Fig. 2 Binding of cP1 with perchlorate salts of divalent cations. Qualitative 1H NMR spectra of cP1 alone and in the presence of perchlorate salts at a molar ratio (metal/cP1) of 2.
The $^1$H NMR spectrum of cyclopeptoid cP1 alone is characterised by a number of benzylic proton resonances between 5 and 6 ppm. This is indicative of the simultaneous presence of cis and trans amides along the backbone. According to literature, for pe side chains, $^1$H NMR resonances between 5.5 and 6 ppm are associated with benzylic protons on cis amides and those around 5 ppm are from benzylic protons on trans amides.17 The average amide cis/trans ratio was estimated to be 75:25 in CD$_3$CN (Fig. 2, 3 and S2†). The addition of a molar excess of metal ion typically resulted in a decrease of the signal around 6 ppm thus indicating an increased population of all-trans conformers. At the saturation point, no further changes of the $^1$H NMR spectrum and the cis/trans ratio were observed. To confirm this, a step-wise titration of strontium perchlorate into cP1 was performed in CD$_3$CN (Fig. 3 and S3†). A simplification of the $^1$H NMR spectrum was observed in the region between 5 and 6 ppm. Specifically, increasing the molar ratio (Sr$^{2+}$/cP1) from 0 to 3 determined a stepwise decrease of the peak around 5.8 ppm with a corresponding increase in the signal around 5 ppm (Fig. 3). This suggested that once the metal complex was formed, the cyclopeptoid cP1 adopted an all-trans conformation. A plot integration of peaks around 5.8 ppm versus the molar ratio clearly indicated that no further changes in the $^1$H NMR spectrum were observed after a 1:1 stoichiometry was reached. The $^1$H NMR spectrum of the Sr$^{2+}$/cP1 complex showed a single set of signals for each α- and β-monomer which therefore was indicative of a discrete conformation with a 3-fold rotational symmetry. NMR temperature studies and 2D experiments enabled full proton assignment of the metal complex (Fig. S4 and Table S1†). Notably, the $^1$H NMR spectrum for the Ca$^{2+}$/cP1 complex is very similar to that of the Sr$^{2+}$/cP1 complex thus also indicative of a 3-fold symmetry (Fig. 2). In contrast the spectra for the Ba$^{2+}$/cP1 and Mg$^{2+}$/cP1 complexes, whilst simpler than that of the peptoid cP1 alone, are distinct from those of Sr$^{2+}$/cP1 and Ca$^{2+}$/cP1 complexes. For all of the metal complexes the cis to trans conversion was complete upon addition of an excess of metal. A cis to trans isomerisation was previously observed in cyclic α-peptoid binding cations.10,18

Circular dichroism. Circular dichroism (CD) was undertaken to further probe and quantify the binding interaction using either electronic CD (ECD) on a benchtop instrument or synchrotron radiation CD (SRCD). CD is an invaluable technique for the characterisation of binding19 due to its sensitivity to subtle conformational changes in the far (160–250 nm) and near (260–340 nm) UV regions, corresponding to electronic transitions of amide and aromatic chromophores, respectively.20 Changes in the CD spectrum of cyclopeptoid cP1 resulting from metal binding can be used for its quantification. To exclude the self-assembly of cyclopeptoid cP1 (i.e. host–host interactions) a preliminary concentration study of cP1 was undertaken (Fig. S5†). In the concentration range between 1 and 500 µM the CD spectra were largely similar, thus excluding intermolecular host–host interactions. The CD spectra upon addition of the metal were therefore indicative of an interaction between the cyclopeptoid cP1 and the metal.

The CD spectrum of cP1 is characterised by two negative maxima at 203 and 222 nm (Fig. 4). Similar CD spectra observed for α-peptoids in CH$_3$CN have been assigned to helical conformations.21 However, due to the cyclic nature of the α,β-peptoid cP1, a helical conformation is unlikely and the spectrum may instead arise from a twisted-like conformation. Based on these observations and the complexity of the $^1$H NMR spectra at room temperature, the CD spectrum of cyclopeptoid cP1 is likely to indicate the presence of multiple conformers rather than one dominant species and the [θ]$_{203}$ nm of $268 \times 10^3$ is therefore taken to be indicative of this. In this case, this dynamic behaviour may be due to the intrinsically flexible nature of the backbone derived from the presence of β-residues and the cis/trans isomerism of tertiary amides.

Consistently with $^1$H NMR data, no binding was observed with sodium and potassium cations as indicated by the lack of
spectral changes in the far UV CD, which was also reported for silver, zinc and iron(II) [Fig. S6†]. This was further confirmed for sodium and zinc by quantitative CD titrations in the far UV region [Fig. S7†]. CD in the near UV and fluorescence spectroscopy also showed no significant spectral changes upon addition of sodium perchlorate to cP1 (Fig. S8†). In contrast, a significant change to the spectral features in the far UV was observed upon addition of perchlorate salts of Ba2+, Ca2+, Fe3+, Mg2+ and Sr2+ thus indicating metal complex formation (Fig. 4, S8 and S9†). Spectroscopic studies by fluorescence and in the near UV by CD showed significant changes upon addition of magnesium perchlorate to cP1 further supporting the metal binding [Fig. S8†].

Nearly identical spectral features were observed for cyclopeptoid cP1 in the presence of Ca2+ or Sr2+. The CD spectra of these metal complexes were characterised by one negative maximum around 208 nm which was red shifted in comparison to the free cyclopeptoid cP1. The CD spectrum in the presence of Ba2+ has two maxima (negative circa 204 nm, positive circa 227 nm). The CD spectrum of the Ba2+ complex has greater [θ] and the negative maximum is only marginally red shifted relative to cP1 and the corresponding Ca2+/Sr2+ complexes. The Mg2+/cP1 complex was characterised by a negative maximum at 198 nm and a positive maximum centred at 222 nm with a component at 216 nm. The CD spectrum in the presence of Fe3+ has two negative maxima (circa 197 nm and circa 210 nm) and shows similarity to the Mg2+ profile based on the blue shift of the negative maxima relative to cP1.

The similarity in appearance of the Sr2+/cP1 and Ca2+/cP1 CD spectra agrees well with the similarity of their 1H NMR spectra. The differences between the CD spectra of these two metal complexes and those with barium and magnesium were also reflected in their 1H NMR spectra suggesting distinct conformational differences. Notably, all metal complexes were found to have increased [θ] relative to cP1 alone which was also consistent with the 1H NMR observation of a conformational shift to population of all-trans isomers i.e. increased ordering of the conformational preference of cP1.

Modelling of metal complexes. In an attempt to understand the possible conformational models for the metal complexes, the semi-empirical modelling method PM6 was utilised as part of MOPAC. In agreement with the 1H NMR data, all metal complexes were modelled in the all-trans forms. Preliminary conformations were assessed following MM2 minimisations in the gas phase before optimisation at the PM6 level in the gas phase. The chosen structures were then optimised by PM6 in acetonitrile using the COSMO solvation model intrinsic to MOPAC. Initial models were constructed from the all-trans form of the cyclopeptoid cP1 by placing the appropriate metal ion at various positions about and in the ring structure followed by minimisation using MM2 and then PM6. The relative stability of the conformations, as judged from heats of formation, for all complexes was principally dependent upon (i) the number of carbonyls binding to the metal cation, (ii) the arrangement of the carbonyls about the metal cation and (iii) the orientations of the ethylphenyl substituents about the ring in that order. A full discussion of all conformations of all ion complexes is beyond the scope of this paper thus Sr2+/cP1 will be used as an exemplar in the following discussion (Fig. S20 and S21†). Structural optimisation in the gas phase gave 4 co-ordinate (4 C=O ligating Sr2+) and 6 co-ordinate species (6 C=O ligating Sr2+) with the 4 co-ordinate systems possessing consistently higher heats of formation (∆Hf ∼ 275 kJ mol−1) than the 6 coordinate. At least two distinct co-ordination patterns involving 6 carbonyl co-ordination were observed, one with ∆Hf values circa 220 kJ mol−1 involving two C=O binding from above and four C=O binding from below the Sr2+ cation, relative to the plane of the ring. In contrast, the conformers approaching C3 symmetric arrangement of C=O about the Sr2+ (3 above and 3 below the plane of the ring) gave significantly lower ∆Hf values, from ∼212 kJ mol−1 down to ∼192 kJ mol−1. Since C3 symmetry was confirmed by 1H NMR system with the lowest ∆Hf was chosen to investigate the arrangement of ethylphenyl groups about the ring. Since the ethylphenyl groups displayed full unhindered rotation about the N-C bond this was an extensive conformational space. To facilitate the search a simple system of varying the arrangements of the ethylphenyl groups about the ring was adopted; in essence systematically changing the C-N-C-C dihedral angles to facilitate various conformational starting structures based upon the relative orientation of the phenyl group ‘above’ (up or u) or ‘below’ (down or d) the plane of the ring relative to the other groups. This analysis revealed a relatively flat potential energy surface about these conformers (Fig. S21†) with minimum energy conformers possessing ∆Hf values from ∼180 to ∼196 kJ mol−1. However the conformations starting from the alternating ethylphenyl (ududud) starting point all gave lower heats of formation than any of the other arrangements. Two minimum energy conformers possessed C2 symmetry, one with ∆Hf = 184.6 kJ mol−1 and one with ∆Hf = 180.1 kJ mol−1. The geometries of these two conformers were then optimised using PM6 with the COSMO solution model inherent in MOPAC. The former conformer did not give a vibrational spectrum without negative frequencies. The latter conformer with further optimisation gave rise to a Sr2+/cP1 complex with three-fold symmetry (C3) about the cyclic axis (Fig. 5 and S10†) with a vibrational spectrum free of negative frequencies. A similar procedure was adopted for the Ba2+ complex giving a Ba2+/cP1 complex with three-fold symmetry (C3) about the cyclic axis. This is in excellent agreement with the NMR spectra (Fig. 2). The average metal–O bond lengths were 2.49 (SD = ±0.05) Å for Sr–O and 2.35 (SD = ±0.04) Å for Ba–O.

In contrast, the lowest energy conformers for both Mg2+/cP1 and Fe3+/cP1 complexes involved only 5 of the 6 carbonyls coordinating to the metal centre. For these complexes the average metal–O bond lengths for the 5 bonding carbonyls were 1.98 (SD = ±0.03) Å for Mg2+/cP1 and 1.91 (SD = ±0.06) Å for Fe3+/cP1 (Fig. 5 and S11a, c†). Consequently, these complexes possessed no symmetry beyond their identity; again, this is in accordance with the results from the 1H NMR for the
Mg\textsuperscript{2+}/cP1 complex (Fig. 2). Replacing the Mg\textsuperscript{2+} and Fe\textsuperscript{3+} ions in the sites of the Ca\textsuperscript{2+} or Sr\textsuperscript{2+} ions of the C\textsubscript{3} symmetrical complexes and optimising the geometry led to stable symmetrical C\textsubscript{3} complexes of these metals (Fig. S11b and d) but their energies were significantly higher (Table S2\textsuperscript{†}) than those illustrated in Fig. 5. Similarly transposing the Ca\textsuperscript{2+}/cP1 and Sr\textsuperscript{2+}/cP1 complexes into those with Mg\textsuperscript{2+} or Fe\textsuperscript{3+} led to rearrangement of the cyclic peptoid cP1 to facilitate 6 coordinate carbonyl binding once again. It is uncertain as to whether the metal complexes with Mg\textsuperscript{2+} and Fe\textsuperscript{3+} shown in Fig. 5 are absolutely the lowest energy conformers due to the range of conformational possibilities available to these complexes however they are significantly lower in energy than any 6 coordinate species found.

The Ba\textsuperscript{2+}/cP1 complex also showed 5 coordinate carbonyl binding but unusually the cation itself was displaced from the plane of the peptoid resulting in close contacts with three of the phenyl groups (Fig. 5 and S12\textsuperscript{†}). Thus average metal–O bond lengths for 3 of the bonding carbonyls were 2.76 (SD = ±0.03) Å and for the other two = 2.56 (SD = ±0.01), with the average C–Ba distance for the three coordinating phenyls being 2.74 (SD = ±0.12) Å. Again, this lack of symmetry was mirrored in the \textsuperscript{1}H NMR and support for the phenyl coordination comes from the appearance of an upfield shifted aromatic resonance at 6.51 ppm.

Overall, the semi-empirical modelling largely agrees with the CD spectra in that the spectra for Ca\textsuperscript{2+}/cP1 and Sr\textsuperscript{2+}/cP1 complexes are virtually identical in line with identical predicted conformations. The maxima for the complexes of Ba\textsuperscript{2+}, Mg\textsuperscript{2+} and Fe\textsuperscript{3+} are distinct from those of Ca\textsuperscript{2+} and Sr\textsuperscript{2+}. However, whereas the negative maxima occur at similar positions for Mg\textsuperscript{2+}/cP1 and Fe\textsuperscript{3+}/cP1 (circa 197 nm) that of the Ba\textsuperscript{2+}/cP1 complex is unique (circa 204 nm) which is in line with the distinct conformation predicted by modelling. Interestingly the CD spectrum of the Fe\textsuperscript{3+}/cP1 complex appears to consist of at least one more distinct peak at circa 210 nm; one possibility for this is that this peak arises from a certain proportion of the complexes in solution adopting the higher energy 6-coordinate C\textsubscript{3} symmetry in equilibrium with the predominately 5-coordinate species.

Quantification of binding
Given the conformational changes observed for cylopeptoid cP1 upon binding and the sensitivity of CD to subtle conformational changes, a stepwise titration of Sr(ClO\textsubscript{4})\textsubscript{2} into cylopeptoid cP1 was performed to give further insight into the binding event (Fig. 6A). The titration showed a gradual increase of the intensity of the CD spectrum for cP1 together with a red shift by 5 nm of the negative maxima around 205 nm and the inversion of the sign for the maximum around 230 nm. An isodichroic point was also observed at 218 nm suggesting a conformational change between two (or more) different conformational states. The differential CD (ΔCD) signal was obtained by subtracting the CD spectrum of cP1 and that of Sr\textsuperscript{2+} from that of the Sr\textsuperscript{2+}/cP1 metal complex for the different molar ratios (Fig. S13\textsuperscript{†}). Using the CD data, a plot of the ΔAbsorbance (ΔA, where ΔA = ΔCD/32 980) at 230 nm as a function of the Sr\textsuperscript{2+} concentration yielded a dissociation constant (K\textsubscript{d}) of 11.9 ± 0.5 µM when fitted to the Hill equation (R\textsuperscript{2} = 0.99) (Fig. 6B).\textsuperscript{22} It is of note that the value of the K\textsubscript{d} was independent of the wavelength selected and the maximum at 231 nm was chosen due to the absence of any significant wavelength shift. The K\textsubscript{d} was of the same order of magnitude as that previously reported for an α-peptoid hexamer with benzylxoyethyl (be) side chains which was shown to bind monovalent cations by quantitative \textsuperscript{1}H NMR.\textsuperscript{10a} This similarity indicated that the α,β-peptoid backbone retained the ability to bind metal ions in the presence of the...
more flexible β-residues and different side chains, albeit with a different cation selectivity. No binding to the monovalent cations investigated was observed for cP1.

A Yoe–Jones plot23 where the differential molar ellipticity at the positive and negative maxima was plotted versus the molar ratio (Sr2+/cP1) gave a stoichiometric point of 1.6 as indicated by the projection of the inflection point on the x-axis (Fig. S13†). This was not consistent with the stoichiometry obtained by 1H NMR and may be rationalised on the basis of the higher sensitivity of CD to subtle conformational changes. Specifically, 1H NMR indicated changes to the ring conformation would not be detected by 1H NMR. By contrast, these would be detected by CD (as an increased signal intensity) and may give rise to the higher stoichiometry observed in comparison to 1H NMR.

Interestingly, when looking at the intensity and wavelength position of the negative maximum at 207, whilst the intensity increases up to a molar ratio of 2, a significant wavelength shift only occurs up to a molar ratio of 1 (Fig. 6A and S14†). Based on these observations and supported by the PM6 calculation showing the 1:1 complex as the most favourable, it is reasonable to assume that the complex is characterised by a 1:1 stoichiometry and any further change in intensity of the CD signal is likely to result from further structural rigidification and an increased population of metal complexes. A similar trend was also observed for the other metal complexes (Fig. S14†) for which similar binding affinity was observed between each metal and cP1 (Table 2 and Fig. S15–S18†).

The Mg2+/cP1 complex gave a Kd of 12 ± 0.2 μM by CD and this was largely comparable with that calculated by isothermal titration calorimetry (ITC),24 which yielded a Kd of the same order of magnitude (2.3 μM) (Fig. 6C). ITC also showed that the binding event was endothermic, as indicated by the positive values for the heat exchange and the enthalpy value (ΔHcal) of 6902 cal mol⁻¹ with an entropic contribution (ΔScal) of 49.2 cal mol⁻¹ K⁻¹. Using ITC, the projection of the mid-point of the titration curve on the x axis provided the stoichiometry of the binding which was found to be close to 1:1 (metal/peptoid) suggesting an approximately equimolecular interaction. This stoichiometry was consistent with that deduced by CD and with the PM6 calculations for the Mg2+/cP1 complex. Unfortunately, quantitative studies for the Mg2+/cP1 complex were hampered by the significant signal overlap (Fig. S19†).

### Binding selectivity

Based on size, there appears to be no correlation in the selectivity of cP1 for divalent over monovalent cations, for example the ionic radius of Na⁺ (not complexed by cP1), is very close to that of Ca²⁺ (1.02 and 1.00 Å, respectively). However, selectivity observed for divalent cations, seems to be at least in part correlated with their greater charge density (Table 1). Amide groups, especially the non-hydrogen-bonded carbonyl groups of peptoid amides, represent polar coordination sites with enhanced divalent ion sensitivity by comparison to peptides and proteins.25 Our results agree well with the observations of Tínés et al. on the selectivity of tetrалactam systems for alkaline-earth over alkaline cations.26 Wipff et al., by working on calix[4]arene amide systems have also observed that complexation of cations is strongly dependent on the nature of the amide groups (primary, secondary and tertiary), with the complexes involving tertiary amides being the most stable.27 Consistent with our results, the higher Sr²⁺/Na⁺ and Ca²⁺/Na⁺ selectivities (ca. 500 and 125, respectively) were measured from calix[4]arene with tertiary amide-based pendant groups. A comparison can also be made between the symmetrical binding mode of Sr²⁺/cP1 and Ca²⁺/cP1 and the Sr₄Symmetry Sr²⁺ complex identified in the solid state for a cyclic α-hexapeptoid bearing side chains.10c It is of note that in this case the coordinating side chains assume a pseudo equatorrial orientation and do not intervene in the binding. Affinity of cyclic α-hexapeptoids (18-membered rings) for alkali metals has also been described, with a peak of selectivity for Na⁺. However, a direct comparison with our results is not fully relevant since the host cyclopeptoids studied carry chelating side chains that were shown to participate in the binding.16d

With regard to the divalent cations, between each other the charge density is not the only criterion to be considered, as their difference in size also plays a significant role in their binding selectivity. The divalent cations with the smallest ionic radii (Mg²⁺, Zn²⁺ and Fe²⁺) bind to cP1 only if the associated charge density is very high. Thus, the hard Mg²⁺ cation binds to cP1 whereas the Zn²⁺ and Fe²⁺ cations having slightly lower charge density do not bind (Table 1). For the other divalent cations tested (Ca²⁺, Sr²⁺ and Ba²⁺), even though their charge density is lower, they have ionic radii ≥1 Å, thus are more adapted to the cavity size, which facilitate their complexation to cP1. Size of the cations also appears to influence the final conformation of the complexes i.e. when the cation radius is too large or too small the binding is 5 coordinate and 6 coordinate only obtainable for a specific range of ionic radii. Despite these observations, the case for the role of the ionic radius and charge density in predicting binding is not clear cut. It would appear that a number of factors involving size, charge density, enthalpy and entropy play a part for all of the species involved, namely the cation, anion, acetonitrile and

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**Table 2. Binding affinity expressed as dissociation constant (Kd) and stoichiometry by CD**

<table>
<thead>
<tr>
<th>Complex</th>
<th>Kd (μM)</th>
<th>Stoichiometry (Mt/cP1)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Sr²⁺/cP1</td>
<td>11.9 ± 0.5</td>
<td>&gt;210 nm, &lt;210 nm</td>
</tr>
<tr>
<td>Mg²⁺/cP1</td>
<td>12 ± 0.2</td>
<td>1.5, 1.5</td>
</tr>
<tr>
<td>Ca²⁺/cP1</td>
<td>8.2 ± 0.3</td>
<td>1.1, 1.04</td>
</tr>
<tr>
<td>Ba²⁺/cP1</td>
<td>12.7 ± 8.6</td>
<td>1.1, 1.2</td>
</tr>
<tr>
<td>Fe²⁺/cP1</td>
<td>1.2 ± 0.9</td>
<td>—, 1.5</td>
</tr>
</tbody>
</table>

---

**Notes:**

1. Using ITC, the projection of the mid-point of the titration curve on the x-axis provided the stoichiometry of Sr²⁺/cP1 as 2/1.
2. Stoichiometry by CD was rationalised on the basis of the differential molar ellipticity. This was not consistent with the stoichiometry obtained by 1H NMR.
3. Stoichiometry by CD was consistent with that deduced by CD and with the PM6 calculations for the Mg²⁺/cP1 complex. Unfortunately, quantitative studies for the Mg²⁺/cP1 complex were hampered by the significant signal overlap (Fig. S19†).

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**References:**

transport in membranes. Alternatively modifications of the environment could be envisaged by cyclic peptoid. Application of this work in a hydrophobic setting as a possible driving force. 

**Conclusion**

We have shown for the first time clear evidence of the selective complexation of divalent over monovalent cations by a 21-membered cyclopeptoid hexamer composed of α- and β-monomers in alternation. This selectivity is different to that generally observed with α-cyclopeptoid hexamers (18-membered rings) which have mainly shown selectivity for monovalent cations. Besides NMR analyses, a detailed analysis of cation binding was undertaken by CD, an invaluable technique in conjunction with other spectroscopic data, in conjunction with other spectroscopic data, to assess subtle conformational changes. From detailed discussion of the combined spectroscopic data, in conjunction with modelling, it is clear that a holistic approach is needed to deduce the reasons for the unique binding behaviour of this cyclic peptoid. Application of this work in a hydrophobic environment could be envisaged e.g. to support selective ion transport in membranes. Alternatively modifications of the side chains could be explored to ascertain if aqueous solubility can be achieved whilst maintaining complexation ability which would allow a wider range of biological applications to be considered.

**Experimental**

**Synthesis and characterisation**

CH₂Cl₂ was distilled under N₂ from CaH₂ and stored over 4 Å molecular sieves. MeOH was distilled under N₂ from CaH₂ and stored over 4 Å molecular sieves. ETOAc, CH₂Cl₂, cyclohexane, and MeOH for column chromatography were distilled before use. DMF and iPr₂NEt were dried over 4 Å molecular sieves. All other solvents and chemicals obtained from commercial sources were used as received. Melting points were determined on a Reichert microscope apparatus and are uncorrected. Specific rotations were measured on a Jasco DIP-370 polarimeter using a 10 cm cell. IR spectra were recorded on a Shimadzu FTIR-8400S spectrometer equipped with a Pike Technologies MIRacle™ ATR and ν are expressed in cm⁻¹. NMR spectra were recorded on a Bruker AC-400 spectrometer or a Bruker AC-500 spectrometer. Chemical shifts are referenced to the residual solvent peak and J values are given in Hz. The following multiplicity abbreviations are used: (s) singlet, (d) doublet, (t) triplet, (q) quartet, (m) multiplet, and (br) broad. Where applicable, assignments were based on COSY, HMBC, HSQC and J-mod-experiments. Flash chromatography was performed with Merck silica gel 60, 40–63 μm. HRMS were recorded on a Micromass Q-Tof Micro (3000 V) apparatus. HPLC analysis was performed on a Waters 590 instrument equipped with an Acclaim® 120 column (C18, 5 μm, 120 Å, 4.6 x 250 mm) and a Waters 484 UV detector.

**Synthesis of cyclic α,β-peptoid cP1**

To a solution of the linear α,β-hexapeptoid¹² (427 mg, 0.40 mmol, 1.0 equiv.) in CH₂Cl₂ (0.10 M) at rt under Ar was added TFA (1 mL per 1 mL CH₂Cl₂) and the resulting mixture was stirred for 2 h at rt. The solvents were evaporated under reduced pressure and the residue was dried in vacuo, yielding the crude termini deprotected peptoid. To a solution of the crude peptoid (5.0 mM) in CH₂Cl₂/DMF (4 : 1) at 0 °C under Ar, enough iPr₂NEt (approx. 5.0 equiv.) was added to turn the mixture slightly basic. HATU (0.48 mmol, 1.2 equiv.) was added and the resulting mixture was stirred for 3 d at rt. The solvents were evaporated under reduced pressure and the residue was taken up in ETOAc (approx. 1/4 of the volume used for cyclization). The organic layer was washed with an equal amount of water and the aqueous layer was extracted with ETOAc (x1). The combined organic layers were dried over MgSO₄, filtered and concentrated under reduced pressure. Flash column chromatography in ETOAc until the impurities had passed followed by change to ETOAc/MeOH 97 : 3 yielded cyclic peptoid cP1 (258 mg, 64%) as a colorless solid: Rf (ETOAc/MeOH 97 : 3) = 0.85; mp = 131–133 °C; [α]D ² = -156.2 (c 0.74, CHCl₃); ¹H NMR (400 MHz, acetone-d₆): δ 7.84–6.70
(0.10H, m, NC\textsubscript{H\textsubscript{6}} × NCH\textsubscript{C}) on a 400 MHz Bruker AC-400 spectrometer or a Bruker AC-500 spectrometer. NMR temperature studies were conducted on a Bruker AC-500 spectrometer equipped with a bespoke temperature control system and fitted with a bespoke thermostated cell holder.\textsuperscript{28} The following parameters were used for SRCD: 4 acquisitions, 10 nm min\textsuperscript{-1} scan speed, 0.1 nm data interval, 1.0 nm spectral band width, 4 s time constant, 260 nm–190 nm scan range. ECD spectra were recorded using a Chirascan Plus instrument (Applied Photophysics Ltd) equipped with a Peltier temperature control system. The following parameters were used for ECD: 4 aquisitions, 1 nm step, 1 s time per point, 1.0 nm spectral band width, 260 nm–190 nm range. The qualitative and quantitative CD titrations were carried out in a stepwise manner by adding small aliquots (3–80 μL) of stock solutions of perchlorate salt of varied concentrations (1, 2 or 3 mM) into a solution of the peptoid of a fixed concentrations (7.5 or 15 μM in the far UV region using 0.4 or 1 cm rectangular pathlength quartz cuvettes from Hellma and 600 μM in the far UV region using a 1 cm cylindrical quartz cuvette). The titration was designed to reach the saturation of the peptoid and a final dilution of the peptoid inferior to 15%. An accurate and precise addition of the small aliquots of the guest, was achieved using positive displacement pipettes. Peptoid spectra (4 accumulations) were recorded at 20 °C in CH\textsubscript{3}CN. Solvent baseline spectra (4 accumulations) were recorded in the same cell at proximal time and averaged. The peptoid spectra were averaged to give a spectrum from which the averaged baseline was subtracted. The resulting baseline corrected spectrum was normalised for pathlength and concentration to give the molar ellipticity (\[θ\], deg cm\textsuperscript{2} dmol\textsuperscript{-1}). CD data from quantitative metal titrations were processed and analysed using a proprietary software developed at the beamline B23 at Diamond Light Source (UK). The software processed the data by taking into account the baseline subtraction, the dilution of the host (peptoid) upon titration and any contributions to the CD deriving from the guest (none for the perchlorate). The binding constant was obtained by fitting of experimental data expressed as ΔAbsorbance (ΔA, where ΔA = ΔCD/32 980) into the Hill equation.\textsuperscript{22}

**Fluorescence spectroscopy**

Fluorescence control system was obtained at 20 °C using a Varian Eclipse instrument equipped with a Peltier temperature control system and a cell with 1 cm path length using a peptoid concentration of 600 μM. Emission spectra were recorded using the following parameters: 250 nm excitation wavelength, 5 nm excitation and emission slit width. Excitation spectra were recorded using the following parameters: 284 nm excitation wavelength, 5 nm excitation and emission slit width. Solvent and metal blanks were also recorded at a proximal time and showed no contribution to the fluorescence in the spectral region investigated.

**Isothermal titration calorimetry (ITC)**

ITC experiments were undertaken using a GE Healthcare microcalorimeter. The reference cell was filled with CH\textsubscript{3}CN, the sample cell (1.8 mL) was filled with a solution of the peptoid at a concentration of 107 μM and titrated with aliquots (2 μL) of a solution of Mg(ClO\textsubscript{4})\textsubscript{2} placed in the syringe (2.5 mM). ITC experiments were undertaken at 20 °C using the
following experimental parameters: 145 injections, 200 seconds spacing between injections, 307 rpm stirring speed and 5 μl s⁻¹ reference power. Titrations of the Mg(ClO₄)₂ into the solvent and of solvent into the peptoids were carried out separately to account for the heats of dissolution and solution respectively and subtracted from the metal – peptoid titration. Each titration was repeated in triplicate to ensure accuracy of the data obtained but data presented herein were not averaged.

**Molecular modelling (PM6)**

Initial models of the complexes were made using ChemDraw (Version 15.0.0.106, Perkin Elmer Informatics) and Chem3D (Version 15.0.0.106, Perkin Elmer Informatics). Initial geometry and conformer searches were conducted using the integral MM2 force field facility in Chem3D.²⁰ Fuller geometry and conformer searches were conducted using the semi-empirical method PM6 in MOPAC2009 via the Chem3D MOPAC interface. Final geometries were first minimised in the gas phase and then minimised and assessed using PM6 with the COSMO model for acetonitrile (dielectric constant set at 37.5).³⁰

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**References**


16 Ba picrate was not used due its limited solubility in CD₃CN.


